A procedure of Arrhenius activation energy determination for salt flotation of particles in the vicinity of one molar salt aqueous solutions

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Abstract: A modified Arrhenius equation, in which thermal energy is replaced with chemical (Gibbs) potential, and a special calculation procedure were used to evaluate the flotation activation energy for salt solutions in the one molar range of concentrations. The proposed formula overcomes the difficulties of calculating the salt flotation activation energy caused by mathematical problems of finding \(1/(\ln(c/c_0))\) for the standard salt concentration \(c_0=1\) M (1 kmol/m\(^3\)) and when the applied salt concentration \(c\) is in the vicinity of 1 M. The salt flotation activation energy was calculated for flotation of copper-bearing carbonaceous shale in the 0.25, 0.50, 1.00 and 2.00 M NaCl solutions performed in a laboratory machine as equal to 1.9 RT for standard salt solution equal to 1 M.

Keywords: activation energy, salt flotation, Arrhenius equation, Gibbs potential

1. Introduction

Flotation in the presence of salt ions is common due to either salty mine or marine waters used in the process (Drzymala, 2007; Laskowski and Castro, 2014). Flotation can also be performed at high, one molar and more, salt concentration or even in salt saturated aqueous solutions (Ratajczak and Drzymala, 2003). The presence of dissolved salts usually increases both the rate and ultimate yield of flotation (Paulson and Pugh, 1996; Pugh et al., 1997; Ozdemir, 2013; Laskowski and Castro, 2015; Zhang, 2015). It was demonstrated in a previous study (Drzymala, 2018) that an Arrhenius-type equation can be used to relate the salt concentration and flotation kinetics, that is to link the value of incentive parameter with the value of a system response parameter, and determine the salt flotation activation energy. The Arrhenius (Arrhenius, 1889) relation, in the logarithmic form, is:

\[
\ln k = \ln A - \frac{E}{RT}
\]

(1)

where \(k\) stands for the rate constant, \(A\) process constant, \(E\) activation energy (J), \(R\) gas constant (J/mol·K), and \(T\) for absolute temperature (K). The Arrhenius relation is used in chemical investigations to find activation energy as a measure of energy barrier of chemical reactions. In the arrheniusan approach \(E\) is the slope of the straight line of a plot of \(\ln k\) versus \(1/T\) while \(RT\) is the thermal energy of reacting species. This procedure of evaluation of chemical reactions can also be used in salt flotation when the thermal energy term \(RT\) is replaced with chemical energy (potential) \(RT \ln (c/c_0)\) (Drzymala, 2018). The arrheniusan relation, adopted to salt flotation, has the form:

\[
\ln k = \ln A - \frac{E}{RT} \ln (c/c_0)
\]

(2)

where \(c\) stands for salt concentration and \(c_0\) is the standard (reference) concentration equal to 1 kmol/m\(^3\).

There is a significant obstacle in determination of the activation energy of a salt flotation with the standard state of the salt concentration equal to 1 M (1 kmol/m\(^3\)), because the applied in flotation salt concentrations are in the order of 1 M, leading to a mathematical problem of calculation of \(1/\ln(c/c_0)=1\). Therefore, in the previous papers on the activation energy in salt flotation (Drzymala,
The data from Fig. 2 were used to create an arrheniusan plot \( \ln k = f(-1/(\ln c - \ln c_0)) \) being a graphical form of Eq. 2. The results of calculation using \( c_0 = 1 \text{ kmol/m}^3 \) are presented in Fig. 3 in the form of three data points.
Fig. 2. Dependence of shale flotation kinetic constant \( k \) on concentration of \( \text{NaCl} \)

Table 1. 1st order kinetic constants, maximum yields and determination coefficients \( R^2 \) of copper-bearing shale flotation performed in \( \text{NaCl} \) aqueous solutions

<table>
<thead>
<tr>
<th>( \text{NaCl, M} )</th>
<th>First order kinetic constant, ( k ), 1/min</th>
<th>Maximum yield, ( y_{\text{max}} ), % (after 30 min of flotation)</th>
<th>Correlation coefficient, ( R^2 ), -</th>
</tr>
</thead>
<tbody>
<tr>
<td>2.00</td>
<td>0.340</td>
<td>96.8</td>
<td>0.998</td>
</tr>
<tr>
<td>1.00</td>
<td>0.238</td>
<td>97.0</td>
<td>0.999</td>
</tr>
<tr>
<td>0.50</td>
<td>0.224</td>
<td>90.0</td>
<td>0.997</td>
</tr>
<tr>
<td>0.25</td>
<td>0.125</td>
<td>82.0</td>
<td>0.997</td>
</tr>
</tbody>
</table>

points indicated by symbol \( \oplus \). The forth data point for \( c_{\text{NaCl}} = 1 \text{ kmol/m}^3 \) could not be displayed due to one-over-zero mathematical problem. As a result, the activation energy of the considered flotation system cannot be determined.

Fig. 3. Relationship between \( \ln k \) and \(-1/(\ln c_{\text{inc}})\) for standard salt concentration \( c_o = 1 \text{ kmol/m}^3 \) (\( \oplus \) points) and for other arbitrarily chosen \( c_o \) values. The mathematical parameters of the approximating solid straight lines provide apparent values of \( E/RT \) and \( A \) while of the dashed line indicate the real values of \( E/RT \) and \( A \) (see also Fig. 4).

To overcome the difficulty in finding \( E \) for systems with reagents concentration near 1 kmol/m\(^3\) and for \( c_o = 1 \text{ kmol/m}^3 \), a special procedure has to be applied. The first step is to plot the relation between \( \ln k \) and \(-1/(\ln c_{\text{inc}})\) for other than 1 M values of \( c_o \). Figure 3 provides such a plot. For instance, for \( c_o=0.0001 \text{ M} \) the relation is \( y = 34.503x + 2.4017 \) meaning that \( E/RT \), that is apparent activation energy, expressed in \( RT \) units, is 34.5RT while apparent \( \ln A \) is 2.4. Assuming other values of \( c_o \), different numerical values of \( A \) and \( E/RT \) are obtained (Fig. 3).
The next step is plotting a relation between $\ln A$ as well as $\ln \frac{E}{RT}$ versus $\ln c_o$. This is shown in Fig. 4.

![Fig. 4. Relationship between $\ln A$ and $\frac{E}{RT}$ for different values of standard NaCl solution $c_o$.](image)

Fig. 4 directly provides the value of $\lim_{c_o \to 1 \text{ kmol/m}^3} \ln A = -1.63$ for $c_o = 1 \text{ M}$. The plot is fully linear. Unfortunately, the linear approximation of $\ln \frac{E}{RT}$ for smaller and greater than 1 M $c_o$ values provides two slightly different value of $E$ at $c_o = 1 \text{ M}$. Therefore, another approximation is needed. We found that plotting $\ln \frac{E}{RT}$ vs $\ln A$, after imposing on the calculating program (Excel) a condition that the lines must cross at the previously determined $\lim_{c_o \to 1 \text{ kmol/m}^3} \ln A$ value -1.63 for $c_o = 1 \text{ M}$, provides activation energy $E$, as shown in Fig. 5.

![Fig. 5. Determination of $\lim_{c_o \to 1 \text{ kmol/m}^3} \ln \left( \frac{E}{RT} \right)$ for $c_o = 1 \text{ M}$ basing on obtained in Fig. 4. $\lim_{c_o \to 1 \text{ kmol/m}^3} \ln A$ value for $c_o = 1 \text{ M}$](image)

According to Fig. 5, the value of $\lim_{c_o \to 1 \text{ kmol/m}^3} \ln \left( \frac{E}{RT} \right)$ is 0.66, meaning that the activation energy $E$ for the investigated shale flotation in the NaCl aqueous solutions is equal to 1.9 RT. It should be noticed that assuming $c_o = 5 \text{ M}$ (Drzymala, 2019) $E/RT = 1.1$ is obtained.

5. Conclusions

The calculation of the Arrhenius activation energy of salt flotation, regulated with the salt concentration in the vicinity of 1 M, is possible provided that a special procedure is applied. It relies on calculation of
activation energy $E$ and flotation rate constant $A$ assuming the reference (standard) salt concentration $c_o$ different from 1M and plotting $\ln A$ vs $\ln c_o$ to find $\lim_{c_o \to 1 \text{kmol/m}^3} \ln A$. Finally, $\ln E/RT$ versus $\ln A$ is plotted and the linear approximation of the data points, forced to cross at the previously determined $\ln A$, provides the seeking value of salt flotation activation energy $E$ expressed in the RT units, which is equal to 1.9 RT, since $\lim_{c_o \to 1 \text{kmol/m}^3} \ln \left( \frac{E}{RT} \right) = 0.66$.

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References


